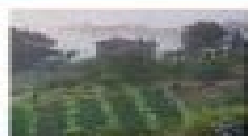


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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar hcn h c n cf4 c f 4 cocl2
c o c l 2 nf3 n f 3 q3

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as bf₃ trigonal planar ch₄ tetrahedral pf₅ trigonal bipyramidal and sf₆

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