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1). Relative atomic mass (ar)

Ans: Average mass of an atom compared with one twelfth of the mass of an atom of carbon 12

2). Electronegativity

Ans: Power of an atom to attract the pair of electrons in a covalent bond

3). Standard enthalpy of combustion

Ans: Enthalpy change when 1 mole of a substance burns completely in oxygen under standard conditions with all reactants and products being in their standard state

Standard enthalpy of formation

Ans: Enthalpy change when 1 mole of a compound is formed from it's elements under standard conditions in their standard states

Mean bond enthalpy

Ans: The average energy needed to break a certain type of bond over a range of compounds

Activation energy

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