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Christian G. Meyer



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web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so 42 aq 2 na aq so 42 aq cucl 2 s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six no 3 aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation 3ba 2 aq 2po 4 3 aq rightarrow ba 3 po 4 2 s

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{OH}^- \rightarrow \text{H}_2\text{O}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web the $\text{Ag}^+ \text{aq}$ and $\text{Cl}^- \text{aq}$ ions become AgCl s but the $\text{Na}^+ \text{aq}$ ions and the $\text{NO}_3^- \text{aq}$ ions stay as $\text{Na}^+ \text{aq}$ ions and $\text{NO}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be $2\text{Cl}^- + \text{Cu}^{2+} \rightarrow 2\text{Cl}^- + \text{Cu}^{2+}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the debye length

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web feb 23 2017 identify the substance oxidized substance reduced reducing agent and reducing agent answer reduction $\text{Ca}^{2+} + 2\text{e}^- \rightarrow \text{Ca}$ oxidation $2\text{K} + 2\text{H}_2\text{O} \rightarrow 2\text{KOH} + \text{H}_2$ the substance oxidized is the reactant that had undergone oxidation k the substance reduced is the reactant that had undergone

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