

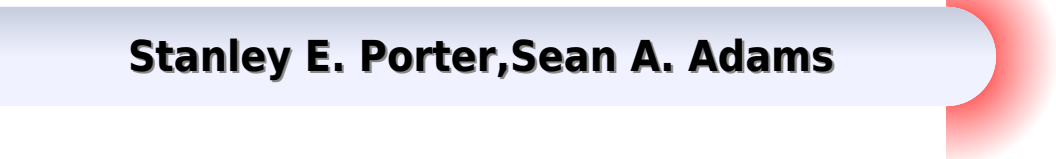
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Basic Language Messages And Meanings

Stanley E. Porter, Sean A. Adams



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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}^{2+}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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