

BASIC CHEMISTRY

CHEMISTRY

- ✓ Medieval Latin "Alchimia"/ Arabic "Alkimya", the chemistry
- ✓ The study of the composition and structure of the changes in matter
- ✓ It is often called the central science because knowledge in this field is a prerequisite for all the other physical and biological sciences.

HISTORICAL DEVELOPMENT OF THE ATOMIC STRUCTURE

ATOM

- ✓ It is the smallest possible unit into which matter can be divided, while still maintaining its properties
- ✓ Building blocks of matter
- ✓ It has protons (+), neutrons (0), and electrons (-)

1. **Democritus** was a greek philosopher who first introduced the concept of atoms. He believed that matter is comprises of tiny indivisible particles known as atomos.
2. **Dalton's Atomic Theory**
 - a. All matter is made up of indivisible particles known as atoms.
 - b. All atoms of the same elements are alike.
 - c. Compounds are formed when atoms of different elements combine in certain proportions.
 - d. atoms are only rearranged but neither created nor destroyed in chemical reactions.
3. **Law of Definite Proportion** – a given compound always contains the same elements in exactly the same proportion by mass.

H_2O - ratio of two hydrogen atom with one oxygen atom only. If it becomes ratio of two hydrogen atom with two oxygen atom, it is not already water.

4. **Law of Multiple Proportions** – when two elements can combine to form more than one type of compound, the masses of one element that combine with a fixed mass of the other elements are in a ratio of small whole numbers.



5. **Law of Conservation of Mass** states that matter can neither be created nor destroyed. Matter is just transformed.
6. The discovery of the different sub-atomic particles:
 - i. **J.J. Thompson** determined the charge to mass ratio of the electron using a cathode ray tube.

Basic Chemistry Student Guide

Karen C. Timberlake



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Table of Contents Basic Chemistry Student Guide

1. Understanding the eBook Basic Chemistry Student Guide
 - The Rise of Digital Reading Basic Chemistry Student Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic Chemistry Student Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic Chemistry Student Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Chemistry Student Guide

- Personalized Recommendations
 - Basic Chemistry Student Guide User Reviews and Ratings
 - Basic Chemistry Student Guide and Bestseller Lists
5. Accessing Basic Chemistry Student Guide Free and Paid eBooks
 - Basic Chemistry Student Guide Public Domain eBooks
 - Basic Chemistry Student Guide eBook Subscription Services
 - Basic Chemistry Student Guide Budget-Friendly Options
 6. Navigating Basic Chemistry Student Guide eBook Formats
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 - Basic Chemistry Student Guide Compatibility with Devices
 - Basic Chemistry Student Guide Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Basic Chemistry Student Guide
 - Highlighting and Note-Taking Basic Chemistry Student Guide
 - Interactive Elements Basic Chemistry Student Guide
 8. Staying Engaged with Basic Chemistry Student Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic Chemistry Student Guide
 9. Balancing eBooks and Physical Books Basic Chemistry Student Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic Chemistry Student Guide
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Basic Chemistry Student Guide
 - Setting Reading Goals Basic Chemistry Student Guide
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Basic Chemistry Student Guide

- Fact-Checking eBook Content of Basic Chemistry Student Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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