

CONCEPT MAP

SOME BASIC CONCEPTS OF CHEMISTRY

Law of Conservation of Mass

- Proposed by the French Chemist Antoine Lavoisier (1789)
- Mass can neither be created nor destroyed in a chemical reaction.
- For any chemical process in a closed system, the mass of the reactants must be equal to the mass of the products.

eg. $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
 $12\text{g} + 32\text{g} = 44\text{g}$
 Oxygen + Hydrogen $\rightarrow$ Water
 $16\text{g} \quad 2\text{g} \quad 18\text{g}$

Law of Definite Proportions

- Proposed by Joseph Proust (1799)
- A chemical compound always contains of the same elements combined together in the same ratio, irrespective of the method of preparation or the source from it is taken.
- In the formation of water compound, the ratio of the mass of hydrogen to the mass of oxygen is always 1 : 8, wherever for the source of water. Thus, if 1 g of water is decomposed, 1 g of hydrogen and 8 g of oxygen are always obtained.

Law of Multiple Proportions

- Proposed by John Dalton (1808)
- When elements combine, they do so in the ratio of small whole numbers. eg. carbon and oxygen react to form CO or CO_2 , but not C_2O_3 .

Concentration Terms

- Mass percent ($\%$) = $\frac{\text{Mass of solute}}{\text{Mass of solution}} \times 100$
- Molarity, $M = \frac{\text{Moles of solute}}{\text{Vol. of solution in l}} \quad (M = \frac{n}{V})$
 $M = (\text{Molarity or Molality}) \times M$
- Molality, $m = \frac{\text{Moles of solute}}{\text{Mass of solvent in kg}}$
- Relative between molarity and molality
 $\frac{1}{m} = \frac{d}{M} \times \frac{M_2}{1000}$
- Molality, $m = \frac{M \times 1000}{M_2 + M \times 1000}$
- Mole fraction, $x_2 = \frac{\text{moles of solute}}{\text{moles of solute} + \text{moles of solvent}}$
 $x_2 = \frac{n_2}{n_1 + n_2}$
- Relative between molarity and mole fraction
 $x_2 = \frac{M \times 1000}{M_2 + M \times 1000}$
 $M = x_2 \times 1000 / (1 - x_2) \times M_2$

Limiting Reagent

The reactant which present in lesser amount and gets consumed and limits the amount of product formed is called limiting reagent.

eg. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 $1\text{g} \quad 71\text{g}$

For every 1g of oxygen, 8g H_2 is required, if H_2 is present less than 8g then it will be limiting reagent.

Avogadro's Law

- Proposed by Avogadro (1811)
- Equal volumes of gases at the same temperature and pressure contain equal number of molecules.

Gay Lussac's Law of Gaseous Volumes

- Proposed by Gay Lussac (1808)
- In gaseous temperature and pressure, the volume of all gaseous reactants and products bear a simple whole number ratio to each other.

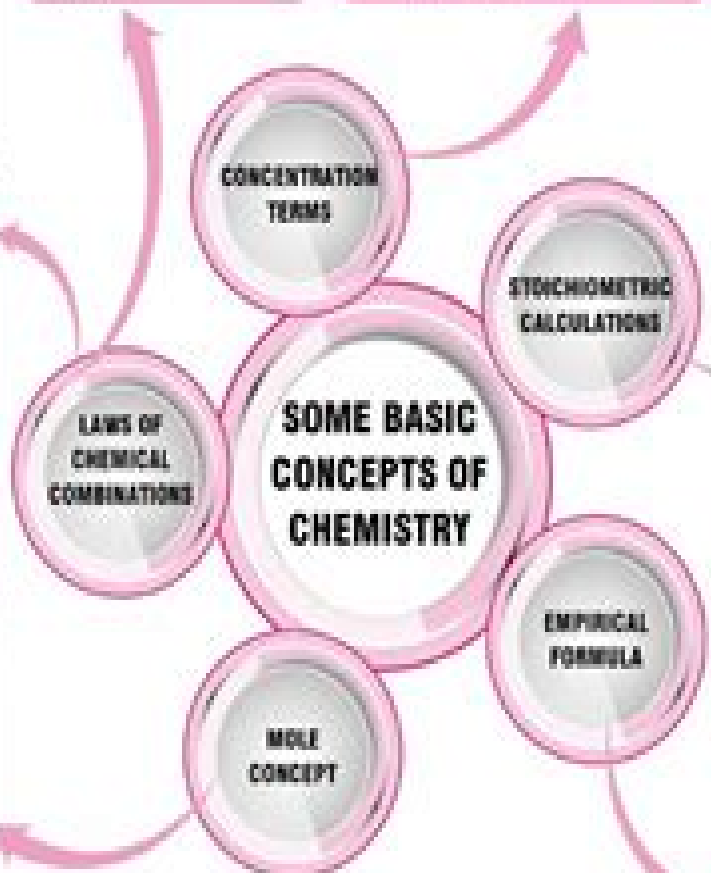
Mole Concept

No. of particles (atoms, molecules, ions) $\rightarrow$ mole $\times 6.022 \times 10^{23}$

Volume (litre) $\xrightarrow{\text{molar gas volume}}$ Mole (mol) $\xleftarrow{\text{molar gas volume}}$ Volume (litre)

Mass (g) $\xrightarrow{\text{molar mass}}$ Mole (mol) $\xleftarrow{\text{molar mass}}$ Mass (g)

Concentration (mol/litre) $\xrightarrow{\text{volume (litre)}}$ Mole (mol) $\xleftarrow{\text{volume (litre)}}$ Concentration (mol/litre)



Mole Calculations

Volume of a gas at STP or NTP $\xrightarrow{V = 22.4\text{ L}} \rightarrow$ Moles of a substance $\xrightarrow{\text{molar mass}} \rightarrow$ Mass of a substance

Mass of a substance $\xrightarrow{\text{molar mass}} \rightarrow$ Moles of a substance $\xrightarrow{V = 22.4\text{ L}} \rightarrow$ Volume of a gas at STP or NTP

Mass of a substance $\xrightarrow{\text{molar mass}} \rightarrow$ Moles of a substance $\xrightarrow{\text{molar mass}} \rightarrow$ Mass of a substance

Molarity and volume of a solution $\xrightarrow{M = \frac{n}{V}} \rightarrow$ Moles of a substance

Moles of a substance $\xrightarrow{\text{molar mass}} \rightarrow$ Mass of a substance

Empirical Formula

Mass % elements $\xrightarrow{\text{Assume 100g sample}} \rightarrow$ Calculate the grams of each element $\xrightarrow{\text{Divide by molar weight}} \rightarrow$ Calculate the moles of each element $\xrightarrow{\text{Calculate the ratio}} \rightarrow$ Assign empirical formula

Molecular Formula

Empirical formula $\xrightarrow{\text{Empirical formula mass}} \rightarrow$ Calculate $n = \frac{\text{Molecular mass}}{\text{Empirical mass}}$ $\rightarrow$ Molecular formula = $n \times$ Empirical formula

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Basic Chemistry:

Understanding Basic Chemistry: The Learner's Approach Kim Seng Chan, Jeanne Tan, 2014-08-13 This book the fourth in a series of Understanding Chemistry books deals with Basic Chemistry Written for students taking either the University of Cambridge O level examinations or the GCSE examinations this textbook covers essential topics under both stipulated chemistry syllabi The book is written in such a way as to guide the reader through the understanding and applications of basic essential chemical concepts by introducing a discourse feature the asking and answering of questions that stimulates coherent thinking and hence elucidates ideas Based on the Socratic Method questions are implanted throughout the book to help facilitate the reader's development in forming logical conclusions of concepts The book helps students to master fundamental chemical concepts in a simple way *Understanding Basic Chemistry Through Problem Solving: The Learner's Approach (Revised Edition)* Kim Seng Chan, Jeanne Tan, 2017-03-02 This book is the revised edition of Understanding Basic Chemistry Through Problem Solving published in 2015 It is in a series of Understanding Chemistry books which deals with Basic Chemistry using the problem solving approach Written for students taking either the university of Cambridge O level examinations or the GCSE examinations this guidebook covers essential topics and concepts under both stipulated chemistry syllabi The book is written in such a way as to guide the reader through the understanding and applications of essential chemical concepts using the problem solving approach The authors have also retained the popular discourse feature from their previous few books Understanding Advanced Physical Inorganic Chemistry Understanding Advanced Organic and Analytical Chemistry Understanding Advanced Chemistry Through Problem Solving and Understanding Basic Chemistry to help the learners better understand and see for themselves how the concepts should be applied during solving problems Based on the Socratic Method questions are implanted throughout the book to help facilitate the reader's development in forming logical conclusions of concepts and the way they are being applied to explain the problems In addition the authors have also included important summaries and concept maps to help the learners to recall remember reinforce and apply the fundamental chemical concepts in a simple way **Basic Chemistry for Water and Wastewater Operators** Darshan Singh Sarai, 2005-06 **MCAT General Chemistry Review 2022-2023** Kaplan Test Prep, 2021-07-06 Kaplan's MCAT General Chemistry Review 2022-2023 offers an expert study plan detailed subject review and hundreds of online and in book practice questions all authored by the experts behind the MCAT prep course that has helped more people get into medical school than all other major courses combined Prepping for the MCAT is a true challenge Kaplan can be your partner along the way offering guidance on where to focus your efforts and how to organize your review This book has been updated to match the AAMC's guidelines precisely no more worrying about whether your MCAT review is comprehensive The Most Practice More than 350 questions in the book and access to even more online more practice than any other MCAT general chemistry book on the market The Best Practice Comprehensive general chemistry subject review is written by top rated

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Notes on General Chemistry Elijah Paddock Harris,1888 **Reconstruction of Wave-Particle Duality and its**

Implications for General Chemistry Textbooks Mansoor Niaz,Cecilia Marcano,2012-04-26 It goes without saying that atomic structure including its dual wave particle nature cannot be demonstrated in the classroom Thus for most science teachers especially those in physics and chemistry the textbook is their key resource and their students core source of information Science education historiography recognizes the role played by the history and philosophy of science in developing the content of our textbooks and with this in mind the authors analyze more than 120 general chemistry textbooks published in the USA based on criteria derived from a historical reconstruction of wave particle duality They come to some revealing conclusions including the fact that very few textbooks discussed issues such as the suggestion by both Einstein and de Broglie and before conclusive experimental evidence was available that wave particle duality existed Other large scale omissions included de Broglie s prescription for observing this duality and the importance of the Davisson Germer experiments as well as the struggle to interpret the experimental data they were collecting Also untouched was the background to the role played by Schrödinger in developing de Broglie s ideas The authors argue that rectifying these deficiencies will arouse students curiosity by giving them the opportunity to engage creatively with the content of science curricula They also assert that it isn t just the experimental data in science that matters but the theoretical insights and unwonted inspirations too In addition the controversies and discrepancies in the theoretical and experimental record are key

drivers in understanding the development of science as we know it today Register of the University of California
University of California (1868-1952),1952 Outlines of General Chemistry Wilhelm Ostwald,1890 *Lithium metal stabilization for next-generation lithium-based batteries: from fundamental chemistry to advanced characterization and effective protection* Yu Yan,Ting Zeng,Sheng Liu,Chaozhu Shu,Ying Zeng,2023-01-11

Lithium Li metal based rechargeable batteries hold significant promise to meet the ever increasing demands for portable electronic devices electric vehicles and grid scale energy storage making them the optimal alternatives for next generation secondary batteries Nevertheless Li metal anodes currently suffer from major drawbacks including safety concerns capacity decay and lifespan degradation which arise from uncontrollable dendrite growth notorious side reactions and infinite volume variation thereby limiting their current practical application Numerous critical endeavors from different perspectives have been dedicated to developing highly stable Li metal anodes Herein a comprehensive overview of Li metal anodes regarding fundamental mechanisms scientific challenges characterization techniques theoretical investigations and advanced strategies is systematically presented First the basic working principles of Li metal based batteries are introduced Specific attention is then paid to the fundamental understanding of and challenges facing Li metal anodes Accordingly advanced characterization approaches and theoretical computations are introduced to understand the fundamental mechanisms of dendrite growth and parasitic reactions Recent key progress in Li anode protection is then comprehensively summarized and categorized to generate an overview of the respective superiorities and limitations of the various strategies Furthermore this review concludes the remaining obstacles and potential research directions for inspiring the innovation of Li metal anodes and endeavors to accomplish the practical application of next generation Li based batteries Outline of Lecture Notes on General Chemistry
John Tappan Stoddard,1885 Science Education in East Asia Myint Swe Khine,2015-09-03

This book presents innovations in teaching and learning science novel approaches to science curriculum cultural and contextual factors in promoting science education and improving the standard and achievement of students in East Asian countries The authors in this book discuss education reform and science curriculum changes and promotion of science and STEM education parental roles and involvement in children s education teacher preparation and professional development and research in science education in the context of international benchmarking tests to measure the knowledge of mathematics and science such as the Trends in Mathematics and Science Study TIMSS and achievement in science mathematics and reading like Programme for International Student Assessment PISA Among the high achieving countries the performance of the students in East Asian countries such as Singapore Taiwan Korea Japan Hong Kong and China Shanghai are notable This book investigates the reasons why students from East Asian countries consistently claim the top places in each and every cycle of those study It brings together prominent science educators and researchers from East Asia to share their experience and findings reflection and vision on emerging trends pedagogical innovations and research informed practices in science education in the region It

provides insights into effective educational strategies and development of science education to international readers

Commencement University of California, Berkeley, 1952 **Basic Chemistry of Life** Milton Toporek, 1968

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