

Construction Estimating Reference Data

Sarviel, Ed

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Construction Estimating Reference Data

Deryl Burch



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web 1 consider the reaction $\text{I}_2\text{O}_5(\text{s}) + 5\text{CO}(\text{g}) \rightarrow \text{I}_2(\text{s}) + 5\text{CO}_2(\text{g})$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

web limiting reactant theoretical and percentage yields key 4 ko2 s 2 h 2 o l 4 koh aq 3 o2 g if a reaction vessel contains 0 mol ko 2 and 0 mol h 2 o what is the limiting reactant

web chem 12 3 limiting reagent and percent yield 4 0 6 reviews whenever quantities of two or more reactants are given in a stoichiometric problem you must identify the limiting reagent

web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂S $\times$ 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1.csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation C₃H₆ + 3O₂ → 3CO + 3H₂O if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
 step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.926 \text{ mol Al}$
 $\frac{5.926 \text{ mol Al}}{4} \times 2 = 2.963 \text{ mol Al}_2\text{O}_3$
 $2.963 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$
 theoretical = 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{}_3\text{PO}_4 + \text{NaNO}_3$ which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 $\times \frac{1}{333.8 \text{ g } \text{I}_2\text{O}_5}$ 1 mol I_2O_5 28 g CO 1 mol CO

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield $\frac{\text{actual yield g}}{\text{theoretical yield g}} \times 100$ label 3 7 3