



Contemporary Mathematics for Business and Consumers

8th Edition

Robert Brechner and George Bergeman

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Robert Brechner



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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 $\text{Na}_2\text{CO}_3 + 4 \text{HCl} \rightarrow 2 \text{NaCl} + \text{CO}_2 + 2 \text{H}_2\text{O}$ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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