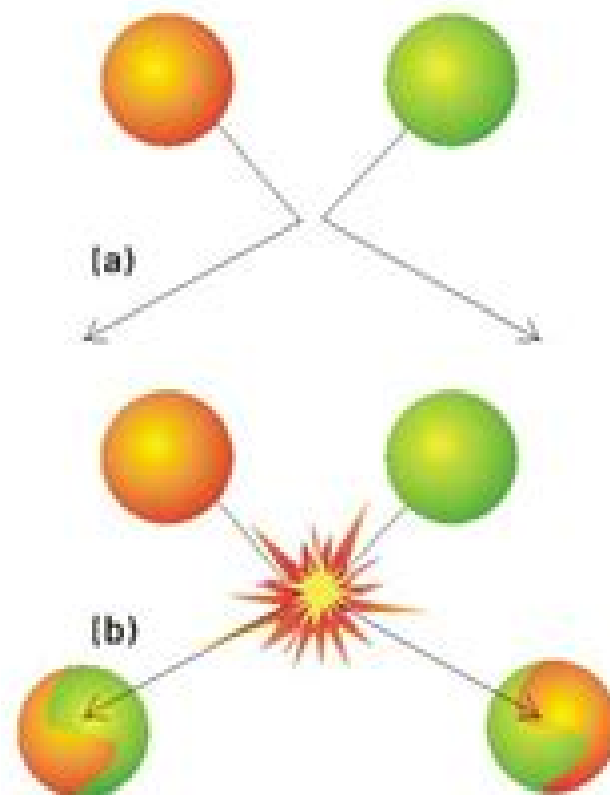


Collision Theory

- In order for a reaction to occur, the particles of the reactant must have enough energy, and must collide at the correct angles (proper orientation).
- The collision theory explains the factors that affect the rate of a reaction.
- The greater the rate of *effective collisions*, the greater the reaction rate is
- The collision theory explains the factors that affect the rate of reaction



Biomolecular Collisions

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