

# Microbiology Study Guide

## Microbial growth

### ***PHYSICAL REQUIREMENTS FOR GROWTH***

- Growth of population: increase in the number of cells
- Requirements are both *physical and chemical*

#### PHYSICAL REQUIREMENTS:

- \*psychrophiles: cold loving microbes
- \*mesophiles: moderate temperature
- \*thermophiles: heat loving
- Minimum growth temperature: the lowest temperature that a species will grow
  - Optimum: temperature it grows best
  - Maximum: highest temperature tolerated for growth
- Most bacteria grow at a pH value between 6.5-7.5
- Most microbes undergo plasmolysis (shrinking) in hypertonic solutions; halophiles (microorganisms that require salt to survive) can tolerate higher salt concentrations

#### CHEMICAL REQUIREMENTS:

- Carbon—all organisms require a carbon source
- \*chemoheterotrophs: use an organic molecule (pure carbon)
- \*autotrophs: use carbon dioxide
- Nitrogen—needed for protein and nucleic acid synthesis. Obtained from decomposition of proteins, from  $\text{NH}_4$  or  $\text{NO}_3$ .
- Oxygen—obligate aerobes (require oxygen); facultative anaerobes (will use oxygen if it is present but do not require it); anaerobes (do not use oxygen) – for growth
- Other: sulfur, phosphorus, trace elements, organic growth factors

#### BIOFILMS

- Microbes adhere to surfaces and accumulate as biofilms on solid surfaces in contact with water
- Most bacteria live in biofilms
- Microbes in biofilms are more resistant to antibiotics than free-swimming microbes

# Basic Microbiology Study Guide

**Karin C. VanMeter, Robert J. Hubert**



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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate  $\text{CuSO}_4$  and 2 0153 grams of iron powder  $\text{Fe}$  the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely

divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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