

BASIC CHEMISTRY

CHEMISTRY

- ✓ Medieval Latin "Alchimia"/ Arabic "Alkimya", the chemistry
- ✓ The study of the composition and structure of the changes in matter
- ✓ It is often called the central science because knowledge in this field is a prerequisite for all the other physical and biological sciences.

HISTORICAL DEVELOPMENT OF THE ATOMIC STRUCTURE

ATOM

- ✓ It is the smallest possible unit into which matter can be divided, while still maintaining its properties
- ✓ Building blocks of matter
- ✓ It has protons (+), neutrons (0), and electrons (-)

1. **Democritus** was a greek philosopher who first introduced the concept of atoms. He believed that matter is comprises of tiny indivisible particles known as atomos.
2. **Dalton's Atomic Theory**
 - a. All matter is made up of indivisible particles known as atoms.
 - b. All atoms of the same elements are alike.
 - c. Compounds are formed when atoms of different elements combine in certain proportions.
 - d. atoms are only rearranged but neither created nor destroyed in chemical reactions.
3. **Law of Definite Proportion** – a given compound always contains the same elements in exactly the same proportion by mass.

H_2O - ratio of two hydrogen atom with one oxygen atom only. If it becomes ratio of two hydrogen atom with two oxygen atom, it is not already water.

4. **Law of Multiple Proportions** – when two elements can combine to form more than one type of compound, the masses of one element that combine with a fixed mass of the other elements are in a ratio of small whole numbers.



5. **Law of Conservation of Mass** states that matter can neither be created nor destroyed. Matter is just transformed.
6. The discovery of the different sub-atomic particles:
 - i. **J.J. Thompson** determined the charge to mass ratio of the electron using a cathode ray tube.

Basic Concepts Of Chemistry Study Guide

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Basic Concepts Of Chemistry Study Guide:

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1994-01-20 This edition has been revamped to improve the logical flow of topics and incorporate more learning aids User friendly it clearly explains the fundamentals of chemical concepts and lessens students intimidation about science Easy to follow example problems help students develop problem solving skills and basic mathematical ideas important in chemistry

Basic Concepts of Chemistry Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance

Sherman Stu Gde Concppts Chem 2ed Sherman, 1980-01-01

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